

Laidler Chemical Kinetics 4th Edition



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Where M stands for concentration in molarity ($\text{mol} \cdot \text{L}^{-1}$), t for time, and k for the reaction rate constant. The half-life of a first order reaction is often expressed as $t_{1/2} = 0.693/k$ (as $\ln 2 = 0.693$). Fractional order. In fractional order reactions, the order is a non-integer, which often indicates a chemical chain reaction or other complex reaction mechanism.

Rate equation - Wikipedia

In chemical kinetics, the overall rate of a reaction is often approximately determined by the slowest step, known as the rate-determining step (RDS) or rate-limiting step. For a given reaction mechanism, the prediction of the corresponding rate equation (for comparison with the experimental rate law) is often simplified by using this approximation of the rate-determining step.

Rate-determining step - Wikipedia

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